

SUBJECTIVE SCORE INSTRUCTOR USE ONLY				
100	90	80	70	60
50	40	30	20	10
9	8	7	6	5
4	3	2	1	0

IMPORTANT

USE NO. 2 PENCIL ONLY

TO USE SUBJECTIVE SCORE FEATURE:

- Mark total possible subjective points
- Only one mark per line on key
- 163 points maximum

EXAMPLE OF STUDENT SCORE:

100	90	80	70	60
50	40	30	20	10
9	8	7	6	5
4	3	2	1	0

• MAKE DARK MARKS

• ERASE COMPLETELY TO CHANGE

• EXAMPLE: (A) (B) (C) (D) (E)

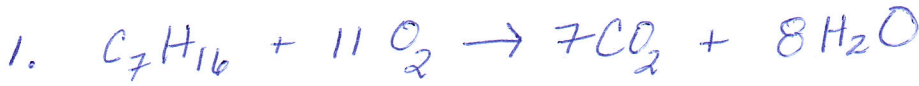
SCANTRON®

NAME	KEY	
SUBJECT		TEST NO. A
DATE		PERIOD

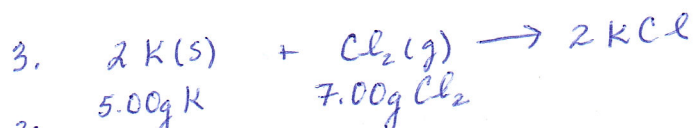
TEST RECORD	
PART 1	
PART 2	
TOTAL	

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PART 1



2. $20.0g C_{10}H_{14}N_2 \times \frac{1 mol C_{10}H_{14}N_2}{162.1g C_{10}H_{14}N_2} \times \frac{2 mol N}{1 mol C_{10}H_{14}N_2} \times \frac{14.01g N}{1 mol N} = 3.45g N$



3a. $5.00g K \times \frac{1 mol K}{39.10g K} = 0.128 mol K \text{ initial}$

$7.00g Cl_2 \times \frac{1 mol Cl_2}{70.90g Cl_2} = 0.0987 mol Cl_2$

$0.128 mol K \times \frac{1 mol Cl_2}{2 mol K} = 0.0640 mol Cl_2, \text{ have } 0.0987 mol Cl_2$

LR = K Cl₂ is excess

3b. Excess Mass

$0.0987 mol Cl_2 \text{ initial}$
 $- 0.0640 mol Cl_2 \text{ used}$

 $0.0347 mol Cl_2 \times \frac{70.90g Cl_2}{1 mol Cl_2} = 2.46g Cl_2 \text{ excess}$

3c. Mass of KCl formed

$0.128 mol K \times \frac{2 mol KCl}{2 mol K} \times \frac{74.60g KCl}{1 mol KCl} = 9.55g KCl$

3d. % yield

$\frac{8.96}{9.55g} \times 100 = 93.8\%$

FEED THIS DIRECTION

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