

TABLE 2.8 Formulas and Names of Some Polyatomic Ions

Ion	Name (Alternate Name in Parentheses)
NH_4^+	ammonium ion
H_3O^+	hydronium ion ^a
OH^-	hydroxide ion
CN^-	cyanide ion
NO_2^-	nitrite ion
NO_3^-	nitrate ion
ClO^- or OCl^-	hypochlorite ion
ClO_2^-	chlorite ion
ClO_3^-	chlorate ion
ClO_4^-	perchlorate ion
MnO_4^-	permanganate ion
$\text{C}_2\text{H}_3\text{O}_2^-$	acetate ion
$\text{C}_2\text{O}_4^{2-}$	oxalate ion
CO_3^{2-}	carbonate ion
HCO_3^-	hydrogen carbonate ion (bicarbonate ion) ^b
SO_3^{2-}	sulfite ion
HSO_3^-	hydrogen sulfite ion (bisulfite ion) ^b
SO_4^{2-}	sulfate ion
HSO_4^-	hydrogen sulfate ion (bisulfate ion) ^b
SCN^-	thiocyanate ion
$\text{S}_2\text{O}_3^{2-}$	thiosulfate ion
CrO_4^{2-}	chromate ion
$\text{Cr}_2\text{O}_7^{2-}$	dichromate ion
PO_4^{3-}	phosphate ion
HPO_4^{2-}	monohydrogen phosphate ion
H_2PO_4^-	dihydrogen phosphate ion

^aYou will encounter this ion only in aqueous solutions.

^bYou will often see and hear the alternate names for these ions.